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Sch4u organic chemistry test answers

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Typically, organic chemistry is taught over the course of two semesters or quarters. The first part introduces key concepts and a few reactions, while the second part elaborates on these concepts and introduces more many reactions. It is the second half, with all of the reactions, that often seems unmanageable. Indeed, the number of reactions discussed would be overwhelming to any student who tried simply to memorize them; however, to the student who understands the underlying mechanisms and principles, the number of reactions covered by the class will seem a lot more manageable. During the first part of the course, concepts such as sterics, resonance, nucleophilicity, electrophilicity, thermodynamic stability, kinetic stability, and electronic inductive effects are introduced. A great deal of organic chemistry can be explained by understanding these seven concepts and how they are related, and thereby a lot of reactions introduced in the second part of the course will make more sense. Consider, as an example, the addition of a Grignard reagent to a ketone. Here, a nucleophile (the Grignard reagent) attacks an electrophile (a positively charged nucleus, i.e. the carbon of the ketone group) to form a carbon-carbon bond. Essentially, this reaction involves an electrostatic attraction between a partially positively-charged carbon (the carbonyl) and a partially negatively-charged carbon (the Grignard). Now, consider the reaction of a Grignard reagent with a nitrile, aldehyde, and ester. These three reactions can be explained by the exact same concepts as the Grignard / ketone reaction! Thus, instead of trying to memorize four reactions, a student, with the proper understanding, can apply one mechanism to predict (without much memorization) the outcome of all of these reactions. Even more good news: other classes of organic chemistry reactions lend themselves to the same process! Consider another example: the dreaded Sn2/Sn1/E1/E2 debate. These are typically the first reactions discussed in any organic chemistry course; however, the principles behind these reactions (nucleophile, good leaving group, steric hindrance) can be applied to the whole course, explaining many reactions later on: for instance, the Sn1 and E1 reactions proceed through a carbocation intermediate. The stability of carbocation intermediates drives many of the organic reactions that are covered in both semesters. Understanding this principle can be as helpful in learning about E1 reactions (first semester) as it can be in learning about allylic cation reactions (second semester). So, to conclude with some practical advice: AVOID MEMORIZATION! Use your time instead to learn and understand the underlying concepts and reaction mechanisms demonstrated by each type of reaction. You can start doing this by taking Varsity Tutors' free Organic Chemistry Practice Tests. Our free Organic Chemistry Practice Tests are each a selection of 10 to 12 questions, which will give you a cross-section of topics from Organic Chemistry. You might think of them as little quizzes, which you can use to hone your skills. Once you realize which topics you need to work on learning, you can focus on understanding the underlying concepts that explain them, and not just memorizing chemical reactions. This will give you a deeper understanding of the material and help you survive, if not thrive, in O-chem! Our completely free Organic Chemistry practice tests are the perfect way to brush up your skills. 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Lesson # Date for this Session Topics to be DIscussed Text Reference and Homework Links 1 Wed Feb 1 text book forms, course outlines, review of grade 11 material, lab safety 2010-09-07 4U Intro Lesson.odp 2010-09-07 Grade 11 Review Work.doc 2010-09-07 First Day Information Page SCH4U.doc 2010-09-07 Lab Notebook Information.doc 4U ISU outline Sept 2k14.doc text p. 1-27 2 Thurs Feb 2 Alkanes (and branched alkanes) (Great resource to view alkane chains here) 2010-09-09 Lesson 2 Hydrocarbon intro lesson.ppt branched alkane worksheet.doc20100914111908.pdf haloalkanes.doc20100914113711.pdf 3 Fri Feb 3 Branched Alkane Practice (modelling lab) NOTE : Branched Alkane Assignment in next column! 2014-09-08 4U Organic Molecule Assignment.doc , p. 28-30, p. 38-45, Q5 p. 42 4 Mon Feb 6 Alkenes and Alkynes alkenew.doc20100914113756.pdf 5 Tues Feb 7 Acetylene Gas Lab 2014-09-06 C2H2 Lab single page.doc 2014-09-11 Lab Assessment Rubric for Reporting and Attaching.docx 6 Wed Feb 8 Cycloalkanes and Aromatics (NOTE: a great practice activity from the University of Calgary can be found here: p.49 - 57, Q 1-3 p.51, Q 4,5 p. 52 cycloalkanes.doc20100914113823.pdf aromatics.doc20100914113847.pdf hydrocarbon reactions.doc20100914113919.pdf 7 Thurs Feb 9 Hydrocarbon reactions p. 45 - 46; Q 11 p. 46 and p.58 - 63 Q 3-7 p. 63. alkane rxn worksheet20100915110324.pdf alkane rxn solutions20100915110353.pd 8 Fri Feb 10 Geometric And Structural Isomerism 9 Mon Feb 13 Alcohols 2010-09-16 Alcohols.ppt p. 64 - 68 Q 1-4 p. 68 10 Tues Feb 14 Aldehydes and Ketones aldehydes and ketones20100917095335.pdf aldehydes and ketones ii20100917095410.pdf p. 69-79, Q 8 p. 78 and Q 4,5 p. 79 Review exercise 11 Wed Feb 15 Carboxylic Acids,Esters and Ethers 2010-09-20 Carboxylic Acids, Esters and Ethers.ppt p. 98 - 155 12 Thurs Feb 16 Amines and Amides 2011-09-23 Amines and Amides.doc 13 Tues Feb 21 Isomerism, Polymers and Interesting Organic Compounds (NOTE: the link provided indicates a number of the different functional groups that we've discussed, as well as others, and ranks them for nomenclature purposes) 2009-09-14 Fats and Proteins and Polymers.ppt 14 Wed Feb 22 Unit I Review Work Period Organic functional groups 15 Jan 09.doc 15 Thurs Feb 23 Unit I Review Take up 2010-09-22 SCH4U Organic Compounds Review.doc 16 Fri Feb 24 Unit I Test organic chemistry unit review solutions20110927103852.pdf (note: document needs to be turned counterclockwise) p. 162 - 166, Q 1-5 p. 166 17 Mon Feb 27 Carleton Field Trip 2009-09-03 4U polymer lab handout.doc Write a concise history of atomic theory prior to 1927 18 Tues Feb 28 Take up Organic Test & Intro Atomic Theory Matter and its Properties, How Quantum Numbers Work and Orbital Shapes Polymer Lab 2010-09-29 The Compton Effect Neutron Discovery.doc 2010-09-29 wave theory and Quantum Numbers.doc quantum numbers20101013110930.pdf quantum numbers orbital shapes20101013111022.pdf 19 Wed Mar 1 Orbital Box Diagrams obd econs and shorthand20101013111108.pdf 20 Thurs Mar 2 Electron Configurations 21 Fri Mar 3 Lab Assignment - Introduce V.S.E.P.R. Theory p.169 - 184, Q 3,5-7 p. 184 2016-03-02 4U Atomic Theory Computer Lab Assignment.doc 22 Mon Mar 6 Atomic Theory Quiz & Lewis Dot Diagrams for simple & Complex Molecules and Ions plus Lewis Diagram work p.186 - 191, Q 3,4 p. 191 p.192-194, Q 6,7 p. 194 Q 8,9 p. 194 23 Tues Mar 7 Take up Quiz & Lewis Dot Diagrams for simple & Complex Molecules and Ions plus Lewis Diagram work 2010-10-06 Bonding Arrangements Pre-QMT.doc p.231 - 235, Q 8-14 p. 235 ions and rules part i20101013111151.pdf rules and bonding ii20101013111221.pdf 24 Wed Mar 8 V.S.E.P.R Theory Read p. 203-207 when finished p. 224 - 226, Q 4,5 p. 227, Q 10-12 p. 229, Q 2-4 p. 230 25 Thurs Mar 9 Finish V.S.E.P.R. Theory and look at hybridized orbitals vsepr part i20101013111249.pdf 26 Fri Mar 10 Determining Net Dipoles for Polar and Non-Polar Molecules p. 242 - 249, Q 10,11 p. 249 vsepr part ii20101013111319.pdf p.251 - 255, Q 6,7 p. 255 vsepr net dipoles i20101013111345.pdf 27 Mon Mar 20 Hybridization and Problems with QMT p. 224 - 226, Q 4,5 p. 227, Q 10-12 p. 229, Q 2-4 p. 230 28 Tues Mar 21 Multiple Bonds, Valence Bond Theory p.235-239 29 Wed Mar 22 Ionic Bonding Theory & Determiniation of Solid properties plus Unit II Review www.chalkbored.com/lessons/chemistry-12/intermolecular-forces.ppt 2010-10-18 Solid Bonding.doc 2010-09-29 QMT downfalls.doc 2012-10-16 4U Atomic Theory Rvw.doc 30 Thurs Mar 23 Unit II Review Take up day 31 Fri Mar 24 Unit II Test Homework: Reporting Enthalpy Changes and Calorimetry p. 298 - 300, Q 1-5 p. 300, p. 313 - 319, Q 1-3 p. 319 32 Mon Mar 27 Thermochemistry Problems with q=mC(delta)T 2010-10-22 q equals mcdt lesson note.doc 2010-10-22 heat capacity worksheet.doc 2010-10-22 Thermochemistry Problems Lengthy.odt 33 Tues Mar 28 Hess' Law and Standard Molar Enthalpies & Hess' Law worksheet p. 322 - 329 and p. 331 - 340, Q 1-3 p. 326 and Q 4,5 p. 329, Q 1 p. 332 and Q 2-5 p. 335 2011-10-25 ME and Hess' Law Projector Font.doc Hess Problems.doc 34 Wed Mar 29 Entropy and Spontaneity, Gibb's Free Energy and State Functions p. 494 - 498, p. 498 - 50-8, Q 2 p. 508 Entropy.doc 2011-03-30 Gibbs Free Energy.doc 35 Thurs Mar 30 Collision Theory p. 360 - 361,p. 365 Q 1,2 p. 361, Q 1-3 p. 364 2010-10-28 Rates & Colln Theory.doc 2010-10-29 Factors Affecting Rates of Reaction.doc 2010-10-28 Collision Theory and EA.doc 36 Fri Mar 31 Hess Law Problems Activity and review work session 37 Mon Apr 3 Thermochemistry Quiz 38 Tues Apr 4 Take up Quiz, Rates and Concentration Part I p. 378 - 391 39 Wed Apr 5 heoretical Rate Law, Order of reaction, Experimental Rate Law p. 367 - p. 370, Q 1,2 p. 371 , p. 372 -377, Q 1-6 p. 377 Rate Law determination.doc 40 Thurs Apr 6 Rates of Reaction Lab - Mg / HCl (LAB) 2010-10-29 Rate Lab Mg Acid.doc 41 Fri Apr 7 Reaction mechanisms p. 392 - 396, Q 1-4 p.397 , p. 399 rate determining step 2 Nov 05.doc Rate Law determination.doc 42 Mon Apr 10 Rates and Concentration Part II - dosage and half life problems half lives and dosage problems 2 Apr 04.doc 43 Tues Apr 11 Unit III Review 2012-11-06 4U Rates and Thermo Review Sheet.DOC 2009-10-30 4U rates and heat review answers.doc 44 Wed Apr 12 Unit III Review Take Up 2012-11-06 4U Rates and Thermo Review Sheet.DOC 2009-10-30 4U rates and heat review answers.doc 45 Thurs Apr 13 Unit III Test p. 424 - 431, Q 1-5 p. 428 p. 441 - 444, Q 1-6 p. 443-445 p. 453 - 456, Q 1-6 p. 457 46 Tues Apr 18 Take up Unit III Test & Intro Equilibrium 2010-11-10 INtro to Equilibrium.doc 2010-11-11 Chemical Equilibrium.doc 47 Wed Apr 19 Type I and Type II & Le Chatalier's Principle type i problems20101112112315.pdf type ii problems20101112112349.pdf 48 Thurs Apr 20 Type III problems & Type IV problems p. 463 - 470, Q 5,6 p.472 a great tutorial about equilibrium intro to eqm20101112112244.pdf 49 Fri Apr 21 Type IV problems 2010-11-12 Calc Kc with Ltd info.doc p. 471 - 476, Q 7,8 p. 476 50 Mon Apr 24 Practice Work on all types of problems E Warren & S Holmes 51 Tues Apr 25 Challenging Questions for Five Types of Problems & discuss answers L Gill & K McKay 2011-11-15 Eqm worksheet problems.doc 52 Wed Apr 26 Quiz and Heterogeneous Equilibrium 2012-04-23 Equilibrium First Half Review.doc p. 476- 480, Q 9,10 p. 480 53 Thurs Apr 27 Take up Quiz and Heterogenous Problems C Melvin & K Stigter p. 54 Fri Apr 28 Ksp problems p. 484 - 486, Q 1-4 p. 486 55 Mon May 1 Predicting Precipitation & TIP Common Ion Problems B Petersen & B Baker 56 Tues May 2 Solving Precipitation Problems when adding solutions together A Morin & M Catt 2012-11-23 4U eqm review.doc p. 488 - 489, Q 5 p. 489 57 Wed May 3 Unit IV review A Eggens & L Darling 2010-11-23 Common Ion Ksp & Pred ppte.doc p. 528 - 532 Q 1-4 p. 486 Q 5-7 p. 489 Q 7-9 p. 492 58 Thurs May 4 Ksp Lab 2012-11-30 SCH4U Ksp Lab.doc Q 7-9 p. 492 59 Fri May 5 Unit IV Test and Recall of Acids and Bases 60 Mon May 8 Takeup Test and Strong Acids / Strong Bases Info D Patey & O McVeigh p. 532-550, Q 1-3 p. 532, Q 4-6 p. 537, Q 8-11 p. 540, Q 12-15 p. 546 Acids and Bases intro.doc 61 Tues May 9 Weak Acid Ionization and Percent Ionization J Beard & J McVeigh p. 551 - 552, p. 554 - 556, Q 3-5 p. 556 2010-12-01 Ka and pka.doc 62 Wed May 10 Percent ionization and ka problems to practice / discuss D Bowman & M Schenck 2009-11-26 percent ionization.doc 63 Thurs May 11 Comparison of pKa pKb and pKw 2010-12-02 pKa and pKb.doc 64 Fri May 12 Acid Base Eqm Quiz 65 Mon May 15 Acid/Base Buffer Systems p. 552 - 554, Q 1,2 p. 554 2010-12-02 Buffer Solutions Part I.doc 66 Tues May 16 Buffer Solutions p. 615 - 620, Q 7,8 p 568; 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